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Question 14 The standard enthalpy of formation ( $\Delta_f H^\circ$ ) at 298 K for methane, CH 4 (g), is -74.8 kJ mol<sup>-1</sup>. The additional information required to determine the average energy for C - H bond formation would be [AIIEE 2006]

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State the Kelvin-Planck and Clausius statements of second law of thermodynamics and prove their equivalence. 6. 0.5 kg of air at 1 bar and 47 °C is compressed in a piston-cylinder assembly to 4 bar and 127°C by doing 5 kJ of work when the surrounding temperature is 27°C.

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Answer: Heat flows into the system, q is +ve. Heat flows out of the system, q is -ve. Work is done on the system, w is +ve. Work is done by the system, w is -ve. Q34: Define First law of thermodynamics. Answer: The total energy of an isolated system remains constant though it may change from one form to another.

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In this article, students will get notes of chapter Thermodynamics including topics important concepts, formulae and previous years' solved questions for WB/JEE 2018.

Revision notes & Important Questions of Thermodynamics  
Solution: N 2 (g) + 3H 2 → 2NH 3 (g)  $\Delta_s = 2 - 4 = - 2 < 0$ . So entropy is negative. Hence option (c) is the answer. 3. If an endothermic reaction is non-spontaneous at freezing point of water and becomes feasible at its boiling point, then. (a)  $\Delta H$  is -ve,  $\Delta S$  is +ve. (b)  $\Delta H$  and  $\Delta S$  both are +ve.

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