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# **Solution Biochemistry**

## **Solution Biochemistry**

In chemistry, a solution is a special type of homogeneous mixture composed of two or more substances. In such a mixture, a solute is a substance dissolved in another substance, known

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as a solvent. The mixing process of a solution happens at a scale where the effects of chemical polarity are involved, resulting in interactions that are specific to solvation.

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Solution, in chemistry, a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility. The term solution is commonly applied to the liquid state of matter, but solutions of gases and solids are possible.

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What parts make up a solution? Solution is a mixture of two components. 1.

Solute: less abundant component, the substance being dissolved 2. Solvent: more abundant component, the

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dissolver Solutions in water are called aqueous (aq). 3.

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The answer for the first part is (A) - this is the carboxyl carbon - it has a single bond to the alpha carbon, a single bond

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to the amine nitrogen, and a double bond to an oxygen, the double bond is what make it  $sp^2$  hybridized. The answer for the second part is (C) - this is the alpha carbon, the bond to hydrogen is implicit and is not drawn.

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A solution is a homogeneous mixture of two substances—that is, it has the same distribution of particles throughout. Technically speaking, a solution consists of a mixture of one or more solutes dissolved in a solvent. The particles of solute and solvent are molecules or ions, with one or more solvent molecules bound to each solute particle.

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But the concentration of the hydrogen ions contributed by the dissociation of water is 100-fold greater than this: 1.0 (



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10 # 7 M. The concentration of the hydrogen ions contributed by the HCl is negligible in comparison. Therefore, the pH of the solution is equal to 7.0. 36. The pH of this solution is 7.0 (see Solution 35).

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