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beyond any doubt that molecules occurred
containing bonding between metal centres
rather than metal interactions, possibly
occurring via the agency of bridging groups
as is $\text{Fe}_2(\text{CO})_9$. The presence of multiple
bonding between metals was recognized,
again by Cotton, in the trimeric ion $[\text{Re}_3\text{Cl}_{12}]^{3-}$. However, as with the iron carbonyl
 $\text{Fe}_2(\text{CO})_9$

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the this is a series of bonds formed between
multiple metal atoms not metal ions and is

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A quadruple bond is a type of chemical
bond between two atoms involving eight
electrons. This bond is an extension of the
more familiar types double bonds and triple
bonds. Stable quadruple bonds are most
common among the transition metals in the
middle of the d-block, such as rhenium,
tungsten, technetium, molybdenum and
chromium. Typically the ligands that

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support quadruple bonds are π -donors, not π -acceptors. The structure of chromium acetate

Quadruple bond - Wikipedia

Covalent radius: the nominal radius of the atoms of an element when covalently bound to other atoms, as deduced from the separation between the atomic nuclei in molecules. In principle, the distance between two atoms that are bound to each other in a molecule (the length of that covalent bond) should equal the sum of their covalent radii.

Atomic radius - Wikipedia

Compounds based primarily on carbon and hydrogen atoms are called organic compounds, and all others are called inorganic compounds. Compounds containing bonds between carbon and a metal are called organometallic compounds.

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Compounds in which components share electrons are known as covalent compounds.

Chemical substance - Wikipedia

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