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Mixed Stoichiometry Problems. 1. Hydrogen and oxygen react under a specific set of conditions to produce water according to the following: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$ A.

Mixed Stoichiometry Problems

Stoichiometry Practice (Selected Answers are given in bold) Mole to

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Mole Problems 1. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ How many moles of hydrogen are needed to completely react with 2.0 moles of nitrogen? 2. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ How many moles of oxygen are produced by the decomposition of 6.0 moles of potassium chlorate? 3. $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$

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ANSWER KEY: Mixed
Stoichiometry

Problems . 1. $2\text{H}_2 + \text{O}_2$
($2\text{H}_2\text{O}$. a). How many
moles of H_2 would be
required to produce 5.0
moles of water? 5.0
moles water. b). What
mass of H_2O is formed
when H_2 reacts with
384 g of O_2 ? 432g H_2 .
2. $\text{H}_2\text{SO}_4 + 2\text{NaOH}$ (
 $\text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$. a).
Balance this equation.
Look above. b).

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Chemistry Worksheet
on Stoichiometry Mixed
Review Assume all
reactions go to
completion. Write the
formula equation,
balance the equations,
and solve the
problems. Draw a
rectangle around the
answer and don't
forget the units.

Methane (CH_4)
combines with oxygen
to form carbon dioxide

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Stoichiometry
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and water. Balanced
equation:

Answers

**Balanced equation:
How many moles of
oxygen are used
with ...**

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Practice Stoichiometry 1

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Answer Key

Play this game to review Chemical Reactions. Using the following equation: $\text{Fe}_2\text{O}_3(\text{s}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{H}_2\text{O}(\text{l})$ How many moles of Fe can be made from 6 moles H_2 ? (This is a one step conversion using mole ratio)

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Stoichiometry Worksheet Answers

Extra Stoichiometry Problems 1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate. a. Write and balance the chemical equation. 2 $\text{AgNO}_3 + \text{BaCl}_2 \rightarrow 2 \text{AgCl} + \text{Ba}(\text{NO}_3)_2$ b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many

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Stoichiometry Problems

Stoichiometry

Worksheet and Key

1.65 mol KClO_3 = 3 mol
 KClO_3 = 3 mol O_2 = 3 mol
2 3.50 mol KCl = 3 mol
 KClO_3 = 0.275 mol Fe
= 3 mol Fe 20 3 = =

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Microsoft Word program, a free Microsoft Word viewer, or a program that can import Word files in order to ...

Classwork and Homework Handouts

MIXED MOLE

PROBLEMS - KEY 1. a)

How many grams are there in 1.5×10^{25} molecules of CO_2 ?

1.110g 1 mol 44.0 g

6.0210molecules

1.510molecules

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CO₁mol 3 23 2

25!=!!!! b) What

volume would the CO₂
occupy at STP? 5.610L

or 560 L 1 mol 22.4 L

6.0210molecules

1.510molecules

CO₁mol 2 23 2

25!=!!!! 2. a) A sample
of NH₃ gas occupies
75.0 liters at STP ...

KEY - CP - Mixed Mole Problems

Mixed Stoichiometry
Problems. View

Worksheet ; Purpose:

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This is the last of the series of four stoichiometry worksheets. This one mixes several different types of problems -- moles to moles, moles to grams, grams to grams, and even some conversions with particles and volume.

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Stoichiometry
review Chemistry.

Worksheet
Answers
What is the first thing
you must do to solve a
stoichiometry problem?

Mixed Stoichiometry Practice | Chemistry Quiz - Quizizz

Mole Conversions and
Stoichiometry Review
Worksheet. 1) Using the
following equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$ How
many grams of sodium
sulfate will be formed if
you start with 200

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grams of sodium hydroxide and you have an excess of sulfuric acid (H_2SO_4)? 2) Using the following equation:

$$Pb(SO_4)_2 + 4 LiNO_3 \rightarrow Pb(NO_3)_4 + 2 Li_2SO_4$$

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B WORKSHEET 12

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Name _____ CHEMISTRY

110 last first. Problem

sets are due within the
first 5 minutes of

lecture on the due date

. Significant figures
must be correct. All set-

ups must be shown for
credit 1] Given the

equation: $2 \text{C}_8\text{H}_{18} +$

$25 \text{O}_2 \rightarrow 16 \text{CO}_2 +$

$18 \text{H}_2\text{O}$ a. How many

moles of oxygen gas

are required to make

8.33 moles of carbon

dioxide?

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WORKSHEET 12

Name last first ...

3 Answer: 5.33 liters

Mixed Stoichiometry

(mixed phases) 9.) The

decomposition reaction

for hydrogen peroxide

is given below. $2\text{H}_2\text{O}_2$

$(\text{l}) \Rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2$

(g) How many grams of

hydrogen peroxide

must be decomposed

to produce 5.5 liters of

oxygen gas at STP?

Answer: 16.7 grams

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10.)

Stoichiometry Problems for Fun and POINTS Name Key facts ...

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