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1982), i.e. the excess Gibbs function is a sum of a local composition term and one or more terms, which take into account the special properties of electrolyte solutions.

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Electrolyte solutions. R. A. Robinson and R. H. Stokes. Academic Press, New York, 1959. xv + 559 pp. \$11.50.

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This classic text, originally published in the 1950s and long since

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out of print, deals primarily with the measurement and interpretation of conductance, chemical potential, and diffusion in solutions of simple electrolytes. Theoretical interpretations are developed in detail, and extensive tables o

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Robinson died in 1979, and I have worked almost entirely on the thermodynamics of non-electrolyte mixtures since 1968, but I am proud to have been associated with him in writing *Electrolyte Solutions*. I. Stokes R H & Robinson R A. Ionic hydration and activity in electrolyte solutions. J. Amer. Chem. Soc. 70:1870-8.1948. (cited 320 times since ...

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The Zdanovskii–Stokes–Robinson (ZSR) relationship [Stokes and Robinson (J. Phys. Chem. 1966, 70, 2126–2130)] enables the solvent content of a liquid mixture to be estimated, for a specified solvent activity, from data for pure solutions of each of the individual solutes. There is an analogous relationship for the activity coefficients of the solutes.

This Week's Citation Classic - Eugene Garfield

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