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### **Answer**

pH and pOH

CONTINUED 1. 0.01 M

HCl  $[H^+] = 0.01M$  pH =

$-\log(0.01) = 2$  2.

0.0010 M NaOH  $[OH^-]$

$= 0.0010M$  pOH =

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 $-\log(0.010) = 3 \rightarrow \text{pH} = 11.3$   
0.050 M NaOH

**pH and pOH**  
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and pOH are a  
measure of how acidic  
or how basic a solution  
is and for most

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### **Page 70 Answers**

pH and pOH The pH of  
a solution indicates  
how acidic or basic that  
solution is. pH range of  
0-7 acidic 7 neutral  
7-14 basic Since

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$[H^+]\cdot[OH^-] = 10^{-14}$  at  $25^\circ C$ , if  $[H^+]$  is known, the  $[OH^-]$  can be calculated and vice versa.  $pH = -\log [H^+]$   
So if  $[H^+] = 10^{-6} M$ ,  $pH = 6$   
 $pOH = -\log [OH^-]$   
So if  $[OH^-] = 10^{-8}$ ,  $pOH = 8$   
Together,  $pH + pOH = 14$ .

## **pH and pOH - Newbury Park High School**

pH/pOH. pH and pOH are a measure of how acidic or how basic a

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solution is and for most solutions, is measured on a scale of 0 to 14. The lower the number, the more acidic the solution is, and the higher the number, the more basic the solution is. Neutral solutions (not acidic or basic) have a pH of 7.

## **pH / pOH - Chemistry Help**

Acids, Bases, pH and pOH . There are several ways to define

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acids and bases, but pH and pOH refer to hydrogen ion concentration and hydroxide ion concentration, respectively. The "p" in pH and pOH stands for "negative logarithm of" and is used to make it easier to work with extremely large or small values.

## **Chemistry Review of pOH Calculations**

Definitions of pH, pOH,



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and the pH scale.

Calculating the pH of a strong acid or base solution. The relationship between acid strength and the pH of a solution.

## **pH, pOH, and the pH scale (article) | Khan Academy**

New York: Prentice Hall. ISBN

0-582-22628-7. pH, pKa, Ka, pKb, and Kb Explained. How to Find pOH in Chemistry. How

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to Calculate the pH of a Weak Acid. What Does pH Stand For? pH Definition and Equation in Chemistry. pH and pKa Relationship: The Henderson-Hasselbalch Equation. Is Negative pH Possible?

## **Here's How to Calculate pH Values - ThoughtCo**

This also is an excellent representation of the concept of pH

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neutrality, where equal concentrations of  $[H^+]$  and  $[OH^-]$  result in having both pH and pOH as 7.

$$pH + pOH = 14.00.$$

$$pH = 14 - pOH.$$

$$pH = 14 - pOH. \text{ Example 1:}$$

What is the pH of a solution that has a pOH of 12.40? Keep in mind that the relationship of pH and pOH equals 14.

$$pH + pOH = 14.00. \quad pOH = 12 \dots$$

## How to Calculate pH

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KEY Chemistry: pH and pOH calculations Part 1: Fill in the missing information in the table below.

pH	[H <sup>3</sup> O <sup>1+</sup> ]	pOH	[OH <sup>1-</sup> ]	ACID or BASE?
3.78	$1.66 \times 10^{-4}$			
10.22	$6.03 \times 10^{-11}$			
M Acid	3.41	$3.89 \times 10^{-4}$		
10.59	$2.57 \times 10^{-11}$			
M Acid	8.81	$1.55 \times 10^{-9}$		
5.19	$6.46 \times 10^{-6}$			
M Base	8.69	$2.04 \times 10^{-9}$		
5.31	$4.88 \times 10^{-6}$			
M Base	8.46	3.47		

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x 10<sup>-9</sup> M 5.54 2.88 x  
10<sup>-6</sup> M Base  
Answers

## **pH and pOH - Ms. Mogck's Classroom**

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