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From above we can see that if we have 12.4 mol  $\text{H}_2$  we need 4.13 mol  $\text{N}_2$ . We don't have that much  $\text{N}_2$  so the .892 mol of  $\text{N}_2$  must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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## Stoichiometry Section 2 Work Section 3 Answer

Modern Chemistry 73  
Stoichiometry

### CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER

Answer the following questions in the space provided. 1. \_\_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of



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## SECTION 2. PROBLEMS

Write the answer on  
the line to the left.

Show all your work in  
the space provided. 1.

The following equation  
represents a laboratory  
preparation for oxygen  
gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$

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furthermore it is not  
directly done,

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Chapter 9 describes how to use mole ratios, molar masses, conversions, limiting reactants, and percent yield to ...

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ANSWER Answer the  
following questions in  
the space provided. 1.

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b The coefficients in a chemical equation represent the (a masses in grams of all reactants and products.

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compounds (Ch 7.3) •  
Reaction stoichiometry  
-The mass  
relationships between  
reactants and products  
in a chemical reaction  
Section 1 Introduction  
to Stoichiometry

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Chapter 9 focuses on  
reaction stoichiometry:  
using a balanced

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chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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## Stoichiometry

### Chapter 9 Answer

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#1 - #18, #31, & #38  
Answers . 38. To

ensure that all  
magnesium is  
converted to  $MgO$ , I  
would use pure  
oxygen, not air, to  
carry out the reaction,  
because  $Mg$  could  
react with  $N_2$  in air to  
form  $Mg_3N_2$ . The pure  
oxygen should be in  
excess. 5. a, 50 mol  $H_2$

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1 SHORT ANSWER  
Answer the following  
questions in the space  
provided. 1. b The  
coefficients in a  
chemical equation  
represent the (a)  
masses in grams of all

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reactants and products. (b) relative number of moles of reactants and products.

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9-3 PROBLEMS Write  
the answer on the line  
to the left. Show all  
your work in the space  
provided. 1. 88% If the

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actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ ; 2.0 mol a.

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Chemical Equations

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82 SHORT ANSWER

Answer the provided a

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9-3 PROBLEMS Write

the answer on the line.

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