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Chemistry - Practice
Problems 1. Fill in the
missing information in
the chart: 2. What is
the nuclear symbol for
a radioactive isotope of
copper with a mass
number of 60? A) Cu B)
Cu C) ^{29}Cu D) Cu E) Cu
3.

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Nuclear Chemistry 8
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Chapter 28 Assignment
& Problem Set Using
the Belt of Stability to
Predict Nuclear
Reactions The best
way to understand
nuclear decay is
determine which
combinations of
protons and neutrons
in a nucleus are stable.
This relationship can
be viewed by plotting
the number of
neutrons (y-axis) vs.
number of protons (x-

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The production of energy in a nuclear reactor can be stopped by pulling out all control rods. A breeder reactor produces more fuel than it uses. The fission products produced in nuclear power plants are not radioactive. An uncontrolled chain reaction led to the nuclear accident in

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Chernobyl, Ukraine.

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Radioactivity~~

- Radioactivity is the process by which nuclei emit particles and rays as they break down.
- The name of the penetrating rays

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emitted by a
radioactive source is
called radiation. •A
radioactive isotope is
an unstable atom
which breaks down on
its own, releasing
energy and/or

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energy levels.

According to the
nuclear shell
model, nucleons exist in
different energy levels,

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or shells, in the nucleus. The numbers of nucleons that represent completed nuclear energy levels—2, 8, 20, 28, 50, 82, and 126—are called magic numbers. NUCLEAR CHEMISTRY 703 FIGURE 22-2 The neutron-proton ratios of stable nuclides clus-

Chapter 25—Nuclear
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Nuclear Chemistry

Nuclear

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Transformations •

Rutherford in 1919 performed the first nuclear transformation.

- The transmutations are sometimes represented by listing in order, the target nucleus, the bombarding particle, the ejecting particle and the product nucleus. • The above equation becomes: ${}_{14}^{27}\text{N} + {}_2^4\text{He} \rightarrow {}_{17}^{14}\text{O} + {}_1^1\text{H}$

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~~Nuclear Reactions •~~

~~Nuclear reactions~~

involve changes in the

nucleus, whereas

chemical reactions

involve the loss, gain,

and sharing of

electrons. • Different

isotopes of the same

element may undergo

very different nuclear

reactions, even though

an element's isotopes

all share the same

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chemical

characteristics.

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Nuclear Chemistry
Chapter 28 Review.

STUDY. PLAY. ... In a
nuclear fission

reaction, a nucleus is
split into 2 smaller,
similar sized nuclei as
well as gives off energy
and 3 neutrons. ...

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induced transmutation
is used to produce a
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Because protons and neutrons reside in the nucleus of atoms, they are called nucleons.

16. There are two forces among the particles within the nucleus. The first, called the electrostatic force, is the force between electrically charged particles. The second force,

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Instructions. Choose your answers to the questions and click

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'Next' to see the next set of questions. You can skip questions if you would like and come back to them ...

~~Chapter 18 Nuclear Chemistry~~

In this lecture I'll teach you about nuclear chemistry. I'll first show you how to determine an element's number of protons, electrons, and neutrons from its atomic symbol. I'll also

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The Nucleus and
Radioactivity Our
journey into the center
of the atom begins
with a brief review. You
learned in Chapter 3
that the protons and
neutrons in each atom

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are found in a tiny,
central

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