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Stoichiometry expresses the quantitative relationship between reactants and products in a chemical equation. Stoichiometric coefficients in a balanced equation indicate molar ratios in that reaction. Stoichiometry allows us to predict certain values, such as the

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percent yield of ...  
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balanced chemical  
equation tell you the  
relative number of  
moles of \_\_\_\_\_ and  
\_\_\_\_\_ in a chemical  
reaction. 9. Why is the

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relative number of  
moles of reactants and  
products the most

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Stoichiometry Atomic  
mass unit and the mole

• amu definition:  $^{12}\text{C}$   
= 12 amu. • 1 amu =  
 $1.6605 \times 10^{-24}$  g •

How many  $^{12}\text{C}$  atoms  
weigh 12 g? •

$6.02 \times 10^{23}$   $^{12}\text{C}$  weigh  
12 g. • Avogadro's  
number • The mole •

#atoms = (1 atom/12

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Stoichiometry  
amu)(1  
amu/1.66x10-24  
g)(12g) = 6.02x10<sup>23</sup>  
12C weigh 12 g

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